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# DESCRIPTION

METHOD FOR ANALYZING COPPER ELECTROPLATING SOLUTION, APPARATUS FOR THE ANALYSIS, AND METHOD FOR FABRICATING SEMICONDUCTOR PRODUCT

## TECHNICAL FIELD

The present invention relates to a method for analyzing a copper electroplating solution used in copper electroplating for filling a copper metal in a via-hole or a trench installed in printed circuit boards, semiconductor package substrates or semiconductor products such as semiconductor substrates, and to an apparatus for the analysis and to a method for producing a semiconductor product. The present application is based on Japanese Patent Application No. 2003-082037 and Japanese Patent Application No. 2003-350544, of which the disclosures are hereby incorporated.

## BACKGROUND ART

In general, additives are used in a copper electroplating solution. Additives include a promoter and a retardant, and when added to a plating solution, then the plating on the bottom of a via-hole is promoted while that outside the via-hole is retarded whereby the via-hole is filled with a plating metal. To attain the phenomenon, the promoter and the retardant used must be well balanced in point of their effects, and the

concentration control of the additive components is important.

One method proposed for the concentration control of additives is CV (cyclic voltammetry) (e.g., see R. Haak, C. Ogden, D. Trench: Plating 64 (4), April 1981). The basis of CV analysis is that, when a voltage is imparted to the inactive electrode in a plating solution and the potential is linearly scanned, then the additive in the solution changes the polarization for copper deposition reaction thereby having an influence on the amount of the deposited copper. The amount of the deposited copper plating thus measured is in proportion to the quantity of electricity having passed through the inactive electrode, and therefore, when the quantity of electricity is measured, then the effect of the additive to change the polarization may be evaluated.

A method that employs this conception is CVS (cyclic voltammetry stripping). Periodically changing the potential applied to an electrode, copper plating on the surface of the electrode and stripping from it is repeated, and the quantity of electricity for stripping is measured as a criterion for the amount of the plating copper. From the ratio of the plating amount with a solution having a known concentration and that with the in-field solution to be analyzed, the additive concentration in the in-field solution is computed.

In spite of the assertion that CV is usable as an additive control apparatus and of the effectiveness of

commercially-available CVS apparatus, there are posed many technical problems. For example, in copper electroplating for filling vias, there is a problem in that the fillability with the plating solution is worsen with the lapse of time after the preparation of the plating bath, even though the additive concentration in the solution is controlled through CVS analysis. This will be because, though the additive component decomposes during plating and therefore its effect reduces, even the component having such a reduced effect may be detected as the additive concentration in CVS analysis.

#### DISCLOSURE OF THE INVENTION

The invention of claim 1 is a method for analyzing a copper electroplating solution containing an additive, which comprises determining the time-dependent potential change at a cathode current density of  $0.1 - 20 \text{ A/dm}^2$  to thereby judge the fillability with the copper electroplating solution.

The invention of claim 2 is the method for analyzing a copper electroplating solution of claim 1, wherein the time-dependent potential change determination is effected while the rotation of the working electrode is controlled in the range of 100 - 7500 rpm.

The invention of claim 3 is the method for analyzing a copper electroplating solution of claim 1 or 2, wherein the fillability is judged by obtaining the potential change speed in the initial stage of electrolysis and the potential convergent

point from the time-dependent potential change curve for a predetermined period of time after the start of the electrolysis.

The invention of claim 4 is the method for analyzing a copper electroplating solution of claims 1 to 3, wherein the fillability is judged by approximating the time-dependent potential change curve for a predetermined period of time after the start of the electrolysis, according to the Boltzmann's function represented by the following numerical formula (1), to thereby obtain the potential change speed  $dx$ :

$$y = \frac{A_1 - A_2}{1 + e^{\frac{x-x_0}{dx}}} + A_2 \quad \dots (1).$$

The invention of claim 5 is an apparatus for analyzing a copper electroplating solution, which comprises using the method for analyzing a copper electroplating solution of any one of claims 1 to 4.

The invention of claim 6 is a method for fabricating a semiconductor product, which comprises using the method for analyzing a copper electroplating solution of any one of claims 1 to 4 and controlling the condition of the solution so as to make the time-dependent change curve profile of the solution near to that of the original solution when it is judged that the condition of the solution is not good and satisfactory fillability could not be obtained.

The invention of claim 7 is a method for analyzing a copper

electroplating solution used in copper electroplating for filling a copper metal in a via-hole or a trench installed in a semiconductor product, which comprises using an electrochemical cell composed of a working electrode (rotary electrode), a reference electrode and a copper electrode (counter electrode) for a copper electroplating solution, electrolyzing the solution with the working electrode (rotary electrode) as a cathode so as to make the cathode current density controlled in the range of  $0.1 - 20 \text{ A/dm}^2$ , determining the time-dependent potential change between the cathode and the reference electrode for a predetermined period of time after the start of the electrolysis, and judging the fillability with the copper electroplating solution from the time-dependent change curve profile.

The invention of claim 8 is the method for analyzing a copper electroplating solution of claim 7, wherein the time-dependent potential change determination is effected while the rotation of the working electrode is controlled in the range of 100 - 7500 rpm.

The invention of claim 9 is the method for analyzing a copper electroplating solution of claim 7 or 8, wherein the fillability is judged by obtaining the potential change speed in the initial stage of electrolysis and the potential convergent point from the time-dependent potential change curve for a predetermined period of time after the start of the electrolysis.

The invention of claim 10 is the method for analyzing a copper electroplating solution of claims 7 to 9, wherein the fillability is judged by approximating the time-dependent potential change curve for a predetermined period of time after the start of the electrolysis, according to the Boltzmann's function represented by the following numerical formula (1), to thereby obtain the potential change speed  $dx$  in the initial stage and the potential convergent point  $A_2$ :

$$y = \frac{A_1 - A_2}{1 + e^{\frac{x - x_0}{dx}}} + A_2 \quad \dots (1).$$

The invention of claim 11 is an apparatus for analyzing a copper electroplating solution, which comprises using the method for analyzing a copper electroplating solution of any one of claims 7 to 10.

The invention of claim 12 is a method for fabricating a semiconductor product, which comprises using the method for analyzing a copper electroplating solution of any one of claims 7 to 10 and controlling the condition of the solution so as to make the time-dependent change curve profile of the solution near to that of the original solution when it is judged that the condition of the solution is not good and satisfactory fillability could not be obtained.

The invention of claim 13 is a method for analyzing a copper electroplating solution containing an additive, which comprises

using an electrochemical cell composed of a working electrode (rotary electrode), a reference electrode and a copper electrode (counter electrode) for a copper electroplating solution, electrolyzing the solution with the working electrode (rotary electrode) as a cathode so as to make the cathode current density controlled in the range of  $0.1 - 20 \text{ A/dm}^2$ , and determining the time-dependent potential change for a predetermined period of time after the start of the electrolysis to thereby judge the uniformity of electrodeposition (film properties and film thickness uniformity) with the solution.

The invention of claim 14 is the method for analyzing a copper electroplating solution of claim 13, wherein the time-dependent potential change determination is effected while the rotation of the working electrode (rotary electrode) is controlled in the range of 100 - 7500 rpm.

The invention of claim 15 is the method for analyzing a copper electroplating solution of claim 13 or 14, wherein the fillability is judged by obtaining the potential change speed in the initial stage of electrolysis and the potential convergent point from the time-dependent potential change curve for a predetermined period of time after the start of the electrolysis.

The invention of claim 16 is the method for analyzing a copper electroplating solution of claims 13 to 15, wherein the uniformity of electrodeposition (film properties and film thickness uniformity) is quantitatively judged by approximating

the time-dependent potential change curve according to the Boltzmann's function represented by the following numerical formula (1), to thereby obtain the potential change speed  $dx$ :

$$y = \frac{A_1 - A_2}{1 + e^{\frac{x-x_0}{dx}}} + A_2 \quad \dots (1).$$

The invention of claim 17 is an apparatus for analyzing a copper electroplating solution, which comprises using the method for analyzing a copper electroplating solution of any one of claims 13 to 16.

The invention of claim 18 is a method for fabricating a semiconductor product, which comprises using the method for analyzing a copper electroplating solution of any one of claims 13 to 16 and controlling the condition of the solution so as to make the time-dependent change curve profile of the solution near to that of the original solution when it is judged that the condition of the solution is not good and satisfactory fillability could not be obtained.

The invention of claim 19 is a method for analyzing a copper electroplating solution used in copper electroplating for filling a copper metal in a via-hole or a trench installed in a semiconductor product, which comprises using an electrochemical cell composed of a working electrode (rotary electrode), a reference electrode and a copper electrode (counter electrode) for a copper electroplating solution, electrolyzing



the solution with the working electrode (rotary electrode) as a cathode so as to make the cathode current density controlled in the range of  $0.1 - 20 \text{ A/dm}^2$ , controlling the rotation of the working electrode (rotary electrode) in two stages falling within a range of  $0 - 7500 \text{ rpm}$ , determining the time-dependent potential change between the cathode and the reference electrode at different rotations, and comparing the time-dependent change curves with each other to thereby judge the fillability with the copper electroplating solution.

The invention of claim 20 is an apparatus for analyzing a copper electroplating solution, which comprises using the method for analyzing a copper electroplating solution of claim 19.

The invention of claim 21 is a method for fabricating a semiconductor product, which comprises using the method for analyzing a copper electroplating solution of claim 19 and controlling the condition of the solution so as to make the time-dependent change curve profile of the solution near to that of the original solution when it is judged that the condition of the solution is not good and satisfactory fillability could not be obtained.

#### BRIEF DESCRIPTION OF THE DRAWINGS

Fig. 1 is a graph showing the result of measurement of the constant-current potential of a plating solution in which the phosphorus-containing copper dipping time is varied.

Fig. 2 is a graph showing the result of measurement of the constant-current potential of a plating solution in which the retardant concentration is standard and the promoter concentration is varied.

Fig. 3 is a graph showing the result of measurement of the constant-current potential of a plating solution in which the promoter concentration is standard and the retardant concentration is varied.

Fig. 4 is a graph showing the change of  $dx$  relative to the dipping time in the result of measurement of the constant-current potential in Fig. 1.

Fig. 5 is a graph showing the change of  $A_2$  relative to the dipping time in the result of measurement of the potential in constant-current electrolysis in Fig. 1.

Fig. 6 is an explanatory view showing one example of an apparatus for analyzing a copper electroplating solution.

Fig. 7 is a graph showing the result of measurement of the constant-current potential of a fresh solution and a deteriorated solution after electrolysis of about 3000 Ahr/liter.

Fig. 8 is a graph showing the result of measurement of the constant-current potential of a plating solution in which the concentration of the additive is varied.

Fig. 9 is a graph showing  $dx$  of the result of measurement of the constant-current potential in Fig. 7.

Fig. 10 is a graph showing the result of measurement of the constant-current potential of a plating solution in which phosphorus-containing copper is not dipped as yet.

Fig. 11 is a graph showing the result of measurement of the constant-current potential of a plating solution in which phosphorus-containing copper was dipped for 24 hours.

Fig. 12 is a graph showing the result of measurement of the constant-current potential of a plating solution in which phosphorus-containing copper was dipped for 96 hours.

Fig. 13 is a graph showing the result of measurement of the constant-current potential of a fresh plating solution and that of the plating solution measured at intervals of 2 months.

Fig. 14 is a graph showing  $\Delta x$  of the result of measurement of the constant-current potential in Fig. 13.

Fig. 15 is a graph showing  $A_2$  of the result of measurement of the constant-current potential in Fig. 13.

Fig. 16 a graph showing the result of measurement of the constant-current potential of a plating solution sampled from a copper plating bath in working for through-hole plating.

Fig. 17 is a graph showing the result of measurement of the constant-current potential of a fresh solution and a deteriorated solution after electrolysis of about 5000 Ahr/liter.

#### BEST MODE FOR CARRYING OUT THE INVENTION

The method for analyzing a copper electroplating solution

of the invention of claims 1 to 6 and claims 7 to 12 is described with reference to one embodiment thereof. Herein investigated is a method capable of differentiating additive-containing copper electroplating solutions from each other as to whether they are good or not good in point of the fillability thereof. The composition of the plating solution comprises 200 g/liter of copper sulfate, 100 g/liter of sulfuric acid, and 50 mg/liter of chloride ion. As the additives to it, used are a promoter containing SPS (bis(3-sulfopropyl)disulfide disodium) and a retardant containing PEG (polyethylene glycol). A suitable amount of a quaternary ammonium salt is added to it as a leveler.

When a phosphorus-containing copper anode is dipped in the copper electroplating solution for a long period of time with no current applied thereto, then the fillability with the solution for plating lowers. This is similar to the phenomenon that may occur in a case where an excess promoter is added to the plating solution. It is believed that a substance having a significant promoting effect may be formed in the solution through the reaction between the metal copper in the surface of the phosphorus-containing copper anode and the additive component. However, when the concentration of the additive component of the plating solution is measured through CVS, then both the promoter and the retardant may have nearly the same concentration before and after dipping a phosphorus-containing copper anode in the solution.

As the solution having good fillability, used is a fresh solution containing a standard amount of an additive and having been just prepared (hereinafter this is referred to as "fresh solution"). As the solution having bad fillability, used is a solution containing a standard amount of an additive in which a phosphorus-containing copper anode was dipped (hereinafter this is referred to as "phosphorus-containing copper-dipped solution"). This is because, when a phosphorus-containing copper is dipped therein, the fillability with the plating solution worsens. The phosphorus-containing copper-dipped solution is prepared by setting a controlled area of 100 cm<sup>2</sup>/liter of a phosphorus-containing copper plate in one liter of the fresh solution, and then statically keeping it as such for 2, 4, 6, 8 and 24 hours.

In the invention of claims 1 to 6 and claims 7 to 12, used is an electrochemical cell composed of a working electrode (rotary electrode), a reference electrode and a copper electrode (counter electrode). In this, the working electrode (rotary electrode) serves as a cathode and the rotation of the working electrode (rotary electrode) is controlled in the range of 100 - 7500 rpm; and an electroplating solution is electrolyzed at a constant current so as to make the cathode current density controlled in the range of 0.1 - 20 A/dm<sup>2</sup>, and the time-dependent potential change between the cathode and the reference electrode is determined. As the reference electrode, used is Ag/AgCl or

Hg/Hg<sub>2</sub>SO<sub>4</sub>, and as the working electrode (rotary electrode), used is a platinum disc electrode, to which, however, the invention should not be limited.

The optimum cathode current density at which the difference between a fresh solution and a phosphorus-containing copper-dipped solution could be known varies depending on the balance of the additive components in the solutions, and therefore it is desirable that the optimum cathode current density is defined individually for each plating solution. In general, the cathode current density may be controlled in the range of 0.1 - 20 A/dm<sup>2</sup>, preferably between 0.1 - 5 A/dm<sup>2</sup>. If it is lower than 0.1 A/dm<sup>2</sup>, then a difference will hardly appear in the measurement results between the solution having good fillability and the solution having poor fillability; but if higher than 5 A/dm<sup>2</sup>, then the potential could hardly become stable.

The rotation of the working electrode during measurement is preferably at least 100 rpm at which the effect of rotation may appear. However, if the rotation is more than 7500 rpm, then it is unfavorable since the mechanical control of the rotation would be difficult. The measurement time may be one within which the potential value may become almost stable, and is preferably shorter. In general, the time may be from 10 to 40 minutes or so.

Fig. 1 shows the result of potential measurement in

constant-current electrolysis of a copper electroplating solution in which the phosphorus-containing copper dipping time is varied. The result is almost the same as the result of potential measurement in the first aspect of the invention. As in Fig. 1, in general, the time-dependent potential change curves greatly change in the initial stage, and tend to reach a constant value with time, but with the prolongation of the phosphorus-containing copper dipping time, the curves separate more from the 0-hour curve. In addition, it is understood that, with the prolongation of the phosphorus-containing copper dipping time, the potential change speed in the initial stage of electrolysis is higher, and the convergent value is shifted more to the positive potential. It is known that the prolongation of the phosphorus-containing copper dipping time worsens the fillability of via-holes, and the time-dependent change curves show a correlation with fillability.

Figs. 2 and 3 show the result of measurement of the constant-current potential of a plating solution in which the additive concentration is varied. In these, the solutions tested are all fresh solutions just prepared by adding additives thereto. Fig. 2 shows the result of measurement of the constant-current potential of a plating solution in which the retardant concentration is standard and the promoter concentration is varied. Fig. 3 shows the result of measurement of the constant-current potential of a plating solution in which

the promoter concentration is standard and the retardant concentration is varied. When the promoter concentration is increased, then the potential change speed in the initial stage is high and the convergent potential is high. When the retardant concentration is increased, then the potential change speed in the initial stage does not change so much and the convergent potential is low. Accordingly, since the potential change speed in the initial stage of electrolysis changes depending on the additive concentration, it is desirable to use solutions having a controlled additive concentration in CVD for fillability evaluation.

The invention of claims 1 to 6 and claims 7 to 12 is for proposing a method and an apparatus for judging fillability in electroplating by determining a time-dependent potential change to obtain time-dependent change curves as in Fig. 1, and using them. One embodiment of the judgment is to use the overall profile of the curves in Fig. 1 to judge the fillability. In Fig. 1, the 0-hour plating solution has the best fillability, and therefore it may be judged that the solutions giving a curve near to the curve profile of the 0-hour plating solution are good.

The characteristics of the curves are further analyzed. Based on the potential change speed in the initial stage of electrolysis and the convergent value as indices, the fillability may be judged. The potential change speed is defined as ( $E_1$



$-E_2)/(t_1 - t_2)$  where  $E_1$  and  $E_2$  each are the potential at a certain time  $t_1$  and  $t_2$  ( $t_1 < t_2$ ), respectively. The convergent value is defined as a potential  $E_3$  at a certain time  $t_3$  ( $t_3 > t_1, t_2$ ) at which the potential change speed is the lowest within the test period. The suitable values of  $t_1$ ,  $t_2$  and  $t_3$  vary depending on the system to which the method is applied, and therefore they may be suitably settled in accordance with the practicability thereof.

In addition, the time-dependent change curve is approximated to the Boltzmann's function represented by the above-mentioned numerical formula (1), to thereby obtain the potential change speed  $dx$  and the potential convergent point  $A_2$ , and the fillability is thereby quantitatively judged. Fig. 4 shows the relationship between  $dx$  obtained from the data in Fig. 1 and the phosphorus-containing copper-dipping time. When the phosphorus-containing copper-dipping time is longer and the potential change speed in the initial stage is higher, then  $dx$  is smaller. When  $dx$  is smaller, then it may be judged that the fillability is bad. Fig. 5 shows the data  $A_2$  obtained similarly. When the phosphorus-containing copper-dipping time is longer and  $A_2$  is nearer to a positive level, then it may be judged that the fillability is bad. For fitting the time-dependent change curve to the Boltzmann's function, usable is a commercially-available software for analysis.

Fig. 6 is a drawing showing one example of an apparatus

for analysis of a copper electroplating solution according to the invention of claim 5 or claim 11. This is the same as one example of the apparatus for analysis of a copper electroplating according to the first aspect of the invention. The number of rotations of the working electrode (rotary electrode) is controlled by the control computer 9 via the controller 10 to be driven only for it. In the rotation-variable stirring unit 1, the rotation should be controllable within a range of 100 - 7500 rpm. Under computer control, an electric current is applied to the electrochemical cell from the voltage/current generator unit 8, and the potential is measured by the multi-meter 7. According to the measurement condition, the constant-current potential measurement is started, and the measured data are displayed in real time and the data collection is started. After the measurement has ended and the data collection has finished, then the time-dependent potential change curve is displayed on the monitor of the control computer, and this may be compared with the previously-recorded curve of a solution having good fillability. In addition, in the control computer 9, the data are computed based on the designated data region, and the potential change speed, the convergent point, and  $dx$  and  $A_2$  in approximation to the Boltzmann's function are displayed.

Using the method and the apparatus of the invention of claims 1 to 6 and claims 7 to 12, the via fillability with a

copper electroplating solution in a semiconductor product can be judged. As a result, when it is judged that the solution condition is not good and satisfactory fillability could not be obtained, then the solution condition must be controlled in order that the time-dependent change curve of the solution could be near to the original condition thereof. For this control, the following methods may be taken.

(1) Control of additive concentration:

The additive components, promoter and retardant are suitably added. The time-dependent potential change curve may be shifted to the positive direction when a promoter is added but to the negative direction when a retardant is added.

(2) Oxidation of copper(I) ion:

When a copper metal such as phosphorus-containing copper anode is dipped in a copper electroplating solution, then the reaction of the additive component with the copper metal produces copper(I) ions, and it is believed that the increase in the copper(I) ion concentration may worsen the fillability. Accordingly, the copper(I) ion is oxidized to copper(II) ion so as to lower the copper(I) ion concentration, whereby the fillability may be improved.

For the oxidation, employable are aeration for aerial oxidation; addition of oxide such as hydrogen peroxide or hypochlorite; and electrolytic oxidation with insoluble anode.

(3) Treatment with activated charcoal:

In particular, when a plating solution is used for a long period of time, then the decomposed product of the additive therein may accumulate in the solution and it may have some influence on the fillability. Accordingly, by removing the decomposed product through treatment with activated charcoal, the plating solution may be restored to its original condition. The treatment with activated charcoal may be effected in a batch process or may be effected in a mode of continuous liquid circulation through an activated charcoal filter. In any case, the treatment may be so controlled that the fillability with the treated solution could be optimized. In this stage, a method of making the time-dependent potential change curve of the treated solution nearer to a case having good fillability is the optimization method.

The method for analyzing a copper electroplating solution according to the invention of claims 13 to 18 is described below. Herein investigated is a method capable of differentiating additive-containing copper electroplating solutions from each other as to whether they are good or not good in point of the deteriorated condition thereof. The composition of the plating solution comprises 75 g/liter of copper sulfate, 190 g/liter of sulfuric acid, and 50 mg/liter of chloride ion. As the additive to it, used is a commercially-available additive to copper plating for printed-wiring boards (Nippon Learonal's Coppergleam ST-901).

As a good solution, used is a fresh solution containing a standard amount of an additive and having been just prepared (hereinafter this is referred to as "fresh solution"). As a bad solution, used is a solution containing a standard amount of an additive and having been used for electroplating operation to a degree of 3000 Ahr/liter or so (hereinafter this is referred to as "deteriorated solution"). This is because, the plating solution is worsened in point of the properties of the film formed with it and of the film thickness uniformity, owing to application of electrolytic load thereto.

In the invention of claims 13 to 18, used is an electrochemical cell composed of a working electrode (rotary electrode), a reference electrode and a copper electrode (counter electrode). In this, the working electrode (rotary electrode) serves as a cathode and the rotation of the working electrode (rotary electrode) is controlled in the range of 100 - 7500 rpm; and, line in the above-mentioned section, an electroplating solution is electrolyzed at a constant current so as to make the cathode current density controlled in the range of 0.1 - 20 A/dm<sup>2</sup>, and the time-dependent potential change between the cathode and the reference electrode is determined. As the reference electrode, used is Ag/AgCl or Hg/Hg<sub>2</sub>SO<sub>4</sub>, and as the rotary electrode, used is a platinum disc electrode, to which, however, the invention should not be limited.

The optimum cathode current density at which the difference

between a fresh solution and a deteriorated solution could be known varies depending on the balance of the additive components in the solutions, and therefore it is desirable that the optimum cathode current density is defined individually for each plating solution. In general, the cathode current density may be controlled in the range of 0.1 - 20 A/dm<sup>2</sup>, preferably between 0.1 - 5 A/dm<sup>2</sup>. If it is lower than 0.1 A/dm<sup>2</sup>, then a difference will hardly appear in the measurement results between the solution having good fillability and the solution having poor fillability; but if higher than 5 A/dm<sup>2</sup>, then the potential could hardly become stable.

The rotation of the working electrode (rotary electrode) during measurement is preferably at least 100 rpm at which the effect of rotation may appear. However, if the rotation is more than 7500 rpm, then it is unfavorable since the mechanical control of the rotation would be difficult. The measurement time may be one within which the potential value may become almost stable, and is preferably shorter. In general, the time may be from 10 to 40 minutes or so.

Fig. 7 shows the result of measurement of the constant-current potential of a fresh solution and a deteriorated solution after electrolysis of about 3000 Ahr/liter. This led to a fact that, with the increase in the electrolysis load after preparation of the plating solution, the potential change speed in the initial stage of electrolysis is higher and the convergent

potential is lower. Specifically, it has been found that, from the potential change speed in the initial stage of electrolysis, the influence of the deteriorated condition of the plating solution on the uniformity of electrodeposition thereof (film properties and film thickness uniformity) can be evaluated.

Fig. 8 shows the result of measurement of the constant-current potential of a plating solution in which the additive concentration is varied. The solutions tested are all fresh solutions just prepared by adding additive thereto. When the additive concentration is increased, then the potential change speed in the initial stage is high and the convergent potential is high. Accordingly, since the potential change speed in the initial stage of electrolysis changes depending on the additive concentration, it is desirable to use solutions having a controlled additive concentration in CVD for fillability evaluation.

It is also desirable that the positions of the three electrodes, working electrode (rotary electrode), reference electrode and copper electrode (counter electrode) are all the time fixed during measurement. The potential in constant-current measurement significantly varies depending on the positional relationship of the three electrodes, and therefore, an apparatus in which the three electrodes must be re-set every time when the test solution is exchanged is unfavorable in the invention. Accordingly, for obtaining data

of good reproducibility, it is desirable that the three electrodes are fixed in the apparatus used and the test solution can be all the time set in the position in which the three electrodes are all dipped in the solution.

Fig. 9 shows dx of the result of measurement of the constant-current potential in Fig. 7. When the potential change speed in the initial stage is higher, then dx is larger. It is judged that the plating solution having a larger dx is one deteriorated to a higher degree.

The apparatus for analyzing a copper electroplating solution according to the invention of claim 17 is the same as one example of the apparatus for analyzing a copper electroplating solution according to the invention of claim 5, shown in Fig. 6, in which the number of rotations of the working electrode (rotary electrode) is controlled by the control computer via the controller 10 to be driven only for it. In the rotation-variable stirring unit 1, the rotation should be controllable within a range of 100 - 7500 rpm. According to the measurement condition, the constant-current potential measurement is started, and the measured data are displayed in real time and the data collection is started. After the measurement has ended and the data collection has finished, then the data are computed by the control computer 9 based on the designated data region, and dx is displayed.

The method for analyzing a copper electroplating solution



according to the invention of claims 19 to 21 is described below. Herein investigated is a method capable of differentiating additive-containing copper electroplating solutions from each other as to whether they are good or not good in point of the fillability thereof. The composition of the plating solution comprises 200 g/liter of copper sulfate (5 hydrate), 100 g/liter of sulfuric acid, and 50 mg/liter of chloride ion. As the additives to it, used are an organic sulfur compound as a promoter and a polyether compound as a retardant, and a suitable amount of a quaternary ammonium salt is added to it as a leveler.

When a phosphorus-containing copper anode is dipped in the copper electroplating solution for a long period of time with no current applied thereto, then the fillability with the solution for plating lowers. This is similar to the phenomenon that may occur in a case where an excess promoter is added to the plating solution. It is believed that a substance having a significant promoting effect may be formed in the solution through the reaction between the metal copper in the surface of the phosphorus-containing copper anode and the additive component. However, when the concentration of the additive component of the plating solution is measured through CVS, then both the promoter and the retardant may have nearly the same concentration before and after dipping a phosphorus-containing copper anode in the solution.

Solutions having the same additive component

concentration are prepared. Like in the first aspect of the invention, a fresh solution containing a standard amount of an additive and having been just prepared (hereinafter this is referred to as "fresh solution"); and the solution in which a phosphorus-containing copper anode was dipped (hereinafter this is referred to as "phosphorus-containing copper-dipped solution") are used for the measurement test. The phosphorus-containing copper-dipped solution is prepared by setting a controlled area of  $100 \text{ cm}^2/\text{liter}$  of a phosphorus-containing copper plate in one liter of the fresh solution, and then statically keeping it as such for 24 and 96 hours.

In the invention of claims 19 to 21, used is an electrochemical cell composed of a working electrode (rotary electrode), a reference electrode and a copper electrode (counter electrode). In this, the working electrode (rotary electrode) serves as a cathode and the rotation of the working electrode (rotary electrode) is controlled to in the range of 0 - 7500 rpm; and an electroplating solution is electrolyzed at a constant current so as to make the cathode current density controlled in the range of  $0.1 - 20 \text{ A/dm}^2$ , and the time-dependent potential change between the cathode and the reference electrode is determined. As the reference electrode, used is Ag/AgCl or Hg/Hg<sub>2</sub>SO<sub>4</sub>, and as the working electrode (rotary electrode), used is a platinum disc electrode, to which, however, the invention

should not be limited.

The optimum cathode current density at which the difference between a fresh solution and a phosphorus-containing copper-dipped solution could be known varies depending on the balance of the additive components in the solutions, and therefore it is desirable that the optimum cathode current density is defined individually for each plating solution. In general, the cathode current density may be controlled in the range of 0.1 - 20 A/dm<sup>2</sup>, preferably between 0.1 - 5 A/dm<sup>2</sup>. If it is lower than 0.1 A/dm<sup>2</sup>, then a difference will hardly appear in the measurement results between the solution having good fillability and the solution having poor fillability; but if higher than 5 A/dm<sup>2</sup>, then the potential could hardly become stable.

The rotation of the working electrode (rotary electrode) during measurement may be suitably selected in two stages of 0 - 7500 rpm. If the rotation is more than 7500 rpm, then it is unfavorable since the mechanical control of the rotation would be difficult. The measurement time may be one within which the potential value may become almost stable, and is preferably shorter. In general, the time may be from 10 to 40 minutes or so.

Fig. 10 shows the result of potential measurement in constant-current electrolysis of a fresh solution at a rotation of the working electrode (rotary electrode) of 0 or 2500 rpm.

Fig. 11 and Fig. 12 each show the result of the same measurement of the solution in which the phosphorus-containing copper dipping time is 24 or 96 hours. From these, it is understood that the time-dependent potential change curve at two-stage rotation significantly varies depending on the phosphorus-containing copper dipping time.

In ordinary via-filling copper electroplating, the flow rate of the plating solution at around the inlet of a via-hole is large, but inside the via-hole, the flow rate is extremely low. Accordingly, it is believed that, in via-filling plating, the condition having good fillability will be such that the deposition reaction could be retarded under a high flow rate condition but the deposition reaction could be promoted under a low flow rate condition. The condition of deposition reaction retardation or promotion may be detected by potential shifting in the negative or positive direction in the potential measurement.

From Figs. 10, 11 and 12, it is understood that, when the dipping time is longer, then the potential difference owing to the rotation difference is larger. Since the potential is in the positive direction (for promotion) at a rotation of 2500 rpm and it is in the negative direction (for retardation) at a rotation of 0 rpm, and this is read as the direction in which the fillability lowers. In fact, it is known that, when the phosphorus-containing copper-dipping time is longer, then the

via-hole fillability lowers, and the time-dependent change curves well describe the fillability behavior.

Fig. 6 is a drawing showing one example of an apparatus for analysis of a copper electroplating solution according to the invention of claim 6. This is the same as one example of the apparatus for analysis of a copper electroplating according to the invention of claim 5.

The number of rotations of the working electrode (rotary electrode) is controlled by the control computer 9 via the controller 10 to be driven only for it. In the rotation-variable stirring unit 1, the rotation should be controllable within a range of 0 - 7500 rpm. According to the measurement condition, the constant-current potential measurement is started, and the measured data are displayed in real time and the data collection is started. After the measurement has ended and the data collection has finished, then the time-dependent potential change curve is displayed on the monitor of the control computer.

Using the method and the apparatus of the invention of claims 19 to 21, the via fillability with a copper electroplating solution in a semiconductor product can be judged. As a result, when it is judged that the solution condition is not good and satisfactory fillability could not be obtained, then the solution condition must be restored to a good one. For this control, the above-mentioned methods of (1) control of additive concentration, (2) oxidation of copper(I) ion, and (3) treatment

with activated charcoal may be taken, like in the invention of claims 7 to 12.

<Example 1>

The invention of claims 1 to 6 is described in detail with reference to the following Examples 1 to 3. The apparatus used has the constitution as in Fig. 6.

100 ml of a plating solution of which the additive concentration was controlled through CVS analysis (plating solution sample 5) was sampled from a via-filling copper plating bath actually in service, and this was put into a beaker and set on a stand 6 for its exclusive use. However, since the plating solution sample changes while left, it was used for the test within 2 days after its sampling.

As the reference electrode 2, used was a two-layer-structured Ag/AgCl reference electrode (the external solution is 10 vol.%  $\text{H}_2\text{SO}_4$ ; the internal solution is 0.1 mol/liter KCl; and Ag/AgCl is in the internal solution, 10 vol.%  $\text{H}_2\text{SO}_4$ ); as the working electrode 3, used was a platinum rotary electrode (electrode area  $4 \text{ cm}^2$ ); and as the counter electrode 4, used was a copper column (diameter 8 mm).

The electric current was so settled that the current density through the platinum rotary electrode could be  $1 \text{ A/dm}^2$ . The voltage range was  $\pm 1 \text{ V}$ ; the current range was 10 mA; the sample distance was 3 s; the rotation of the platinum rotary electrode was 2500 rpm. The test time was 40 minutes. With

that, the constant-current potential measurement was started.

After the measurement and the data collection ended, the time-dependent potential change curve was automatically approximated to the above-mentioned Boltzmann's function, and  $dx$  was thereby computed. Based on the level of  $dx$ , the via fillability with the solution sample was judged.

<Example 2>

From a via-filling copper plating bath actually in service, the plating solution was sampled at intervals of 2 months, and subjected to constant-current potential measurement under the same condition as in Example 1. The additive concentration of the solution was controlled to be constant through CVS analysis.

Fig. 13 shows the result of measurement of the constant-current potential of a fresh solution having the same additive concentration as that of the above solution, and that of the plating solution sampled at intervals of 2 months. Fig. 14 shows  $dx$  of the result of measurement of the constant-current potential in Fig. 13. From Figs. 13 and 14, it is understood that, even though the additive concentration was controlled to be constant through CVS analysis, the fillability still lowered.

<Example 3>

A fresh solution was prepared and divided into two. Nothing was put into one; and a phosphorus-containing copper plate with a controlled area of  $100 \text{ cm}^2/\text{liter}$  was put into the other and statically left as it was for 48 hours. Using the

two solutions, a via-hole having a diameter of 100  $\mu\text{m}$  and a depth of 75  $\mu\text{m}$  was plated. With the former solution, the via-hole was filled, but not with the latter.

Both solutions were subjected to CVS analysis to determine the promoter concentration and the retardant concentration therein, but the difference between the two was only on a level of error. From this, it may be said that the additive concentration control through CVS analysis is not enough to judge the via fillability.

<Example 4>

The invention of claims 7 to 12 is described in detail with reference to the following Examples 4 and 5. The apparatus used has the constitution as in Fig. 6.

A copper electroplating solution having a composition of 200 g/liter of copper sulfate (5 hydrate), 100 g/liter of sulfuric acid and 50 mg/liter of chloride ion was prepared, to which were added a suitable amount of an organic sulfur compound, a polyether compound and a quaternary ammonium compound. The fresh solution was divided into two. One was kept as such; and a phosphorus-containing copper plate with a controlled area of 100  $\text{cm}^2$ /liter was put into the other and statically left as it was for 24 hours. Using the two solutions, a substrate with a via-hole having a diameter of 100  $\mu\text{m}$  and a depth of 75  $\mu\text{m}$ , the surface of which was made electroconductive as having a 1- $\mu\text{m}$  layer made in a mode of electroless copper plating thereon, was



plated. With the former solution, the via-hole was filled, but not with the latter. Both solutions were subjected to CVS analysis to determine the promoter concentration and the retardant concentration therein, but there was little difference between the two. From this, it may be said that the additive concentration control through CVS analysis is not enough to judge the via fillability.

The solution was analyzed according to the method of the invention. 100 ml of the plating solution was collected, put into a beaker, and set on a stand 6 for its exclusive use. As the reference electrode 2, used was Ag/AgCl; as the working electrode (rotary electrode) 3, used was a platinum disc electrode (electrode area  $4 \pi \text{mm}^2$ ); and as the counter electrode 4, used was a copper column (diameter 8 mm).

The electric current was so settled that the current density through the platinum rotary electrode could be  $1 \text{ A/dm}^2$ . The voltage range was  $\pm 1 \text{ V}$ ; the current range was  $10 \text{ mA}$ ; the sample distance was  $3 \text{ s}$ ; the rotation of the platinum rotary electrode was  $2500 \text{ rpm}$ . The test time was 40 minutes. With that, the constant-current potential measurement was started.

After the measurement and the data collection ended, the time-dependent potential change curve was automatically displayed on the monitor panel of a computer. The displayed curves are the same as the time-dependent change curves (0 hr, 24 hrs) in Fig. 1, and they significantly differ from each other

in point of their profile. From this, it is confirmed that the difference in the fillability between plating solutions can be detected according to the invention, which, however, could not be detected in CVS analysis.

Further, the time-dependent change curve was approximated to the Boltzmann's function, and  $dx$  and  $A_2$  were thereby computed. The data are shown in Table 1. It has been confirmed that the solutions of worsened fillability can be judged owing to the reduction in the value  $dx$  thereof and to the shifting of the value  $A_2$  thereof in the positive direction.

Table 1

	0 hr	24 hrs
$dx$	586	46
$A_2$ (V vs. Ag/AgCl)	-0.082	-0.036

#### <Example 5>

From a via-filling copper plating bath actually in service, the plating solution was sampled at intervals of 2 months, and subjected to constant-current potential measurement under the same condition as in Example 4. The additive concentration of the solution was controlled to be constant through CVS analysis. For comparison, a fresh solution of which the additive concentration was controlled in the same manner was subjected to the same measurement.

Fig. 13 shows the result of measurement of the constant-current potential of a fresh solution having the same

additive concentration as that of the above solution, and that of the plating solution sampled at intervals of 2 months. The results were the same as those in Example 2. Figs. 14 and 15 show  $dx$  and  $A_2$ , respectively, of the result of the potential measurement in constant-current electrolysis of Fig. 13. As in Fig. 13, the time-dependent potential change curve profile varies even though the additive concentration was controlled to be constant through CVS analysis. This confirms that the method of the invention gives information which could not be obtained in CVS analysis.

Regarding the via-hole fillability with the solution, the solution sample collected in 6 months gave a hollowed profile at around the mouth of the via-hole and its fillability was insufficient. As in Figs. 7 and 9, the data  $dx$  and  $A_2$  vary. By detecting and controlling these data, the fillability of the solution could be evaluated.

#### <Example 6>

The invention of claims 13 to 18 is described in detail with reference to the following Examples 6 and 7. The apparatus used has the constitution as in Fig. 6.

100 ml of a plating solution of which the additive concentration was controlled through CVS analysis was sampled from a copper plating bath actually in service for through-hole plating, and this was put into a beaker and set on a stand 6 for its exclusive use. The plating solution sample was used

for the test within 2 days after its sampling.

The solution was analyzed according to the method of the invention. 100 ml of the plating solution was sampled, put into a beaker and set on a stand 6 for its exclusive use. As the reference electrode 2, used was Ag/AgCl reference electrode; as the working electrode (rotary electrode) 3, used was a platinum disc electrode (electrode area  $4 \text{ mm}^2$ ); and as the counter electrode 4, used was a copper column (diameter 8 mm).

The electric current was so settled that the current density through the platinum rotary electrode could be  $1 \text{ A/dm}^2$ . The voltage range was  $\pm 1 \text{ V}$ ; the current range was  $10 \text{ mA}$ ; the sample distance was  $3 \text{ s}$ ; the rotation of the platinum rotary electrode was  $2500 \text{ rpm}$ . The test time was 20 minutes. With that, the constant-current potential measurement was started. Fig. 16 shows the result of the measurement of the constant-current potential of the above plating solution.

After the measurement and the data collection ended, the time-dependent potential change curve was automatically approximated to the above-mentioned Boltzmann's function, and  $dx$  was thereby computed. Based on the level of  $dx$ , the deteriorated condition of the solution sample was judged.

#### <Example 7>

A fresh solution was prepared and divided into two. One was used for electrolysis at about  $5000 \text{ Ahr/liter}$ , and the resulting deteriorated solution was used. The two solutions

were subjected to constant-current potential measurement under the same condition as in Example 6. The additive concentration of the two solutions was controlled to be constant through CVS analysis.

Fig. 17 shows the result of the constant-current potential measurement of the fresh solution having the same additive concentration as above, and that of the deteriorated solution after electrolysis at about 5000 Ahr/liter. It is understood that, even though the additive concentration is controlled to be constant through CVS analysis, the worsened condition of the plating solution became further worsened.

Using the two solutions, a via-hole having a diameter of 100  $\mu\text{m}$  and a depth of 75  $\mu\text{m}$  was plated. The solutions were investigated for the film properties and the throwing power. The former had a throwing power of about 73 %, a tensile strength of 312  $\text{N/mm}^2$ , a ductility of 19 %, and a hardness of 121 Hv; while the latter had a throwing power of about 51 %, a tensile strength of 302  $\text{N/mm}^2$ , a ductility of 15.6 %, and a hardness of 120 Hv. In point of the throwing power and the ductility, the deterioration of the latter was found.

#### <Example 8>

The invention of claims 19 to 21 is described in detail with reference to the following Example 8. The apparatus used has the constitution as in Fig. 6.

A copper electroplating solution having a composition of

200 g/liter of cupper sulfate (5 hydrate), 100 g/liter of sulfuric acid and 50 mg/liter of chloride ion was prepared, to which were added a suitable amount of an organic sulfur compound, a polyether compound and a quaternary ammonium compound. The fresh solution was divided into two. One was kept as such; and a phosphorus-containing copper plate with a controlled area of  $100 \text{ cm}^2$ /liter was put into the other and statically left as it was for 96 hours.

Using the two solutions, a substrate with a via-hole having a diameter of  $100 \text{ }\mu\text{m}$  and a depth of  $75 \text{ }\mu\text{m}$ , the surface of which was made electroconductive as having a  $1\text{-}\mu\text{m}$  layer made in a mode of electroless copper plating thereon, was plated. With the former solution, the via-hole was filled, but not with the latter. Both solutions were subjected to CVS analysis to determine the promoter concentration and the retardant concentration therein, but there was little difference between the two. From this, it may be said that the additive concentration control through CVS analysis is not enough to judge the via fillability.

The solution was analyzed according to the method of the invention. 100 ml of the plating solution was collected, put into a beaker, and set on a stand 6 for its exclusive use. As the reference electrode 2, used was Ag/AgCl; as the working electrode (rotary electrode) 3, used was a platinum disc electrode (electrode area  $4 \text{ mm}^2$ ); and as the counter electrode 4, used was a copper column (diameter 8 mm).

The electric current was so settled that the current density through the platinum rotary electrode could be  $1 \text{ A/dm}^2$ . The voltage range was  $\pm 1 \text{ V}$ ; the current range was  $10 \text{ mA}$ ; the sample distance was  $3 \text{ s}$ ; the rotation of the platinum rotary electrode was first  $0 \text{ rpm}$ , and then it was  $2500 \text{ rpm}$  after the solution sample exchange. The test time was 40 minutes.

Fig. 10 shows the result of the test with the fresh solution; and Fig. 12 shows the result of the test with the solution in which the phosphorus-containing copper was dipped for 96 hours. From this, it is confirmed that the difference in the fillability between plating solutions can be detected according to the invention, which, however, could not be detected in CVS analysis.

As described in detail hereinabove with reference to Examples, the invention of claims 1 to 6 has solved the problem with plating solutions in that the fillability with aged plating solutions worsens even though the additive concentration therein is controlled through CVS analysis. According to the analysis of the invention, it has become possible to judge the plating solutions having the same additive concentration in CVS analysis as to whether their fillability is good or not good. When the fillability of a plating solution is judged not good as a result of the analysis of the invention, then the additive may be removed from the solution or an additional additive may be added to the solution, and the invention has made it possible to more accurately control a plating solution to a higher level.

The invention of claims 7 to 12 has solved the problem with copper electroplating solutions to be used for filling a via-hole with a copper metal in a mode of copper electroplating, in that the via fillability worsens owing to the time-dependent change of the plating solution. According to the analysis of the invention, it has become possible to judge the plating solutions having the same additive concentration in CVS analysis as to whether their fillability is good or not good. When the fillability of a plating solution is judged not good as a result of the analysis of the invention, then the additive concentration of the solution may be controlled or the solution may be oxidized or treated with activated charcoal, and the invention has made it possible to more accurately control a plating solution to a higher level.

The invention of claims 13 to 18 has solved the problem with plating solutions in that the film properties and the uniformity worsen even though the additive concentration in the solutions is controlled through CVS analysis. According to the analysis of the invention, it has become possible to judge the plating solutions having the same additive concentration in CVS analysis as to whether their fillability is good or not good. When the uniformity of electrodeposition (film properties and film thickness uniformity) of a plating solution is judged not good as a result of the analysis of the invention, then the additive may be removed from the solution or an additional additive may



be added to the solution, and the invention has made it possible to more accurately control a plating solution to a higher level.

The invention of claims 19 to 21 has solved the problem with copper electroplating solutions to be used for filling a via-hole with a copper metal in a mode of copper electroplating, in that the via fillability worsens owing to the time-dependent change of the plating solution. According to the analysis of the invention, it has become possible to judge the plating solutions having the same additive concentration in CVS analysis as to whether their fillability is good or not good. When the fillability of a plating solution is judged not good as a result of the analysis of the invention, then the additive concentration of the solution may be controlled or the solution may be oxidized or treated with activated charcoal in order that the time-dependent change curve profile of the solution may be nearer to the original profile thereof, and the invention has made it possible to more accurately control a plating solution to a higher level.

#### INDUSTRIAL APPLICABILITY

Application examples of the invention include copper electroplating solutions used for filling a copper metal in a via-hole or a trench installed in printed circuit boards, semiconductor package substrates or semiconductor products such as semiconductor substrates, and the invention provides a method for analyzing the solutions that require higher level control

and provides an apparatus for the analysis.